



Acta Crystallographica Section E

## Structure Reports

Online

ISSN 1600-5368

## 5-(3,5-Difluorophenyl)-1-(4-fluorophenyl)-3-trifluoromethyl-1H-pyrazole

Karikere Ekanna Manoj Kumar,<sup>a</sup> Parameshwar Adimoole Suchetan,<sup>a</sup> Bandrehalli Siddagangaiah Palakshamurthy,<sup>b</sup> Shankar Madan Kumar,<sup>c</sup> Neratur Krishnappagowda Lokanath<sup>c</sup> and Swamy Sreenivasa<sup>a\*</sup>

<sup>a</sup>Department of Studies and Research in Chemistry, U.C.S., Tumkur University, Tumkur, Karnataka 572 103, India, <sup>b</sup>Department of Studies and Research in Physics, U.C.S., Tumkur University, Tumkur, Karnataka 572 103, India, and <sup>c</sup>Department of Studies in Physics, University of Mysore, Manasagangotri, Mysore, India  
Correspondence e-mail: drsreenivasa@yahoo.co.in

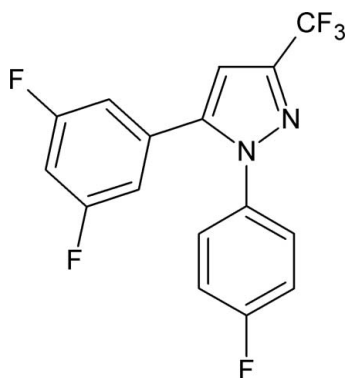
Received 25 November 2013; accepted 1 December 2013

Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.004$  Å;  $R$  factor = 0.054;  $wR$  factor = 0.181; data-to-parameter ratio = 10.1.

In the title compound,  $\text{C}_{16}\text{H}_8\text{F}_6\text{N}_2$ , the dihedral angle between the pyrazole and difluorobenzene rings is  $50.30(13)^\circ$ , while those between the pyrazole and fluorobenzene rings and between the difluorobenzene and fluorobenzene rings are  $38.56(13)$  and  $53.50(11)^\circ$ , respectively. Aromatic  $\pi-\pi$  stacking interactions between adjacent difluorobenzene rings [centroid-centroid separation =  $3.6082(11)$  Å] link the molecules into dimers parallel to  $[21\bar{2}]$ .

## Related literature

For background to pyrazole derivatives and their uses, see: Ramaiah *et al.* (1999). For a similar structure, see: Sreenivasa *et al.* (2013).



## Experimental

## Crystal data

$\text{C}_{16}\text{H}_8\text{F}_6\text{N}_2$   
 $M_r = 342.24$   
 Triclinic,  $P\bar{1}$   
 $a = 7.2535(3)$  Å  
 $b = 8.6686(4)$  Å  
 $c = 11.7690(5)$  Å  
 $\alpha = 70.909(1)^\circ$   
 $\beta = 80.139(1)^\circ$   
 $\gamma = 88.077(1)^\circ$   
 $V = 688.78(5)$  Å<sup>3</sup>  
 $Z = 2$   
 Cu  $K\alpha$  radiation  
 $\mu = 1.39$  mm<sup>-1</sup>  
 $T = 293$  K  
 $0.39 \times 0.35 \times 0.29$  mm

## Data collection

Bruker APEXII CCD diffractometer  
 Absorption correction: multi-scan (SADABS; Bruker, 2009)  
 $T_{\min} = 0.611$ ,  $T_{\max} = 0.669$   
 7069 measured reflections  
 2181 independent reflections  
 2040 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.042$

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.054$   
 $wR(F^2) = 0.181$   
 $S = 1.11$   
 2181 reflections  
 217 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.45$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.48$  e Å<sup>-3</sup>

Data collection: APEX2 (Bruker, 2009); cell refinement: SAINT-Plus (Bruker, 2009); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: Mercury (Macrae *et al.*, 2008); software used to prepare material for publication: SHELXL97.

The authors acknowledge the IOE X-ray diffractometer Facility, University of Mysore, Mysore, for the data collection.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: WM2788).

## References

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## supporting information

*Acta Cryst.* (2014). E70, o19 [https://doi.org/10.1107/S1600536813032650]

**5-(3,5-Difluorophenyl)-1-(4-fluorophenyl)-3-trifluoromethyl-1*H*-pyrazole**

**Karikere Ekanna Manoj Kumar, Parameshwar Adimoole Suchetan, Bandrehalli Siddagangaiah Palakshamurthy, Shankar Madan Kumar, Neratur Krishnappagowda Lokanath and Swamy Sreenivasa**

**S1. Comment**

The pyrazole entity is an important moiety in numerous natural and synthetic compounds and in medicinal chemistry (see, for example: Ramaiah et al., 1999). As part of our studies in this area, the title compound, C<sub>16</sub>H<sub>8</sub>F<sub>6</sub>N<sub>2</sub>, was synthesized and its crystal structure determined.

In the title compound, the dihedral angle between the pyrazole and the difluorobenzene rings is 50.30 (13)°, while those between the pyrazole and the fluorobenzene rings and the difluorobenzene and the fluorobenzene rings are, respectively, 38.56 (13)° and 53.50 (11)°. Compared to these values, the dihedral angles between the pyrazole-benzoic acid ring, pyrazole-fluorobenzene ring and fluorobenzene-benzoic acid ring in the structure of the related compound 2-[5-(2-fluorophenyl)-3-isobutyl-1*H*-pyrazol-1-yl]benzoic acid (Sreenivasa et al., 2013) are 53.1 (1)°, 52.1 (1)° and 62.1 (1)°, respectively. Aromatic  $\pi$ - $\pi$  stacking interactions between the difluorobenzene rings in the title structure [centroid-to-centroid separation = 3.6082 (11) Å] links the molecules parallel to [21 $\bar{2}$ ] in the crystal structure (Fig 2).

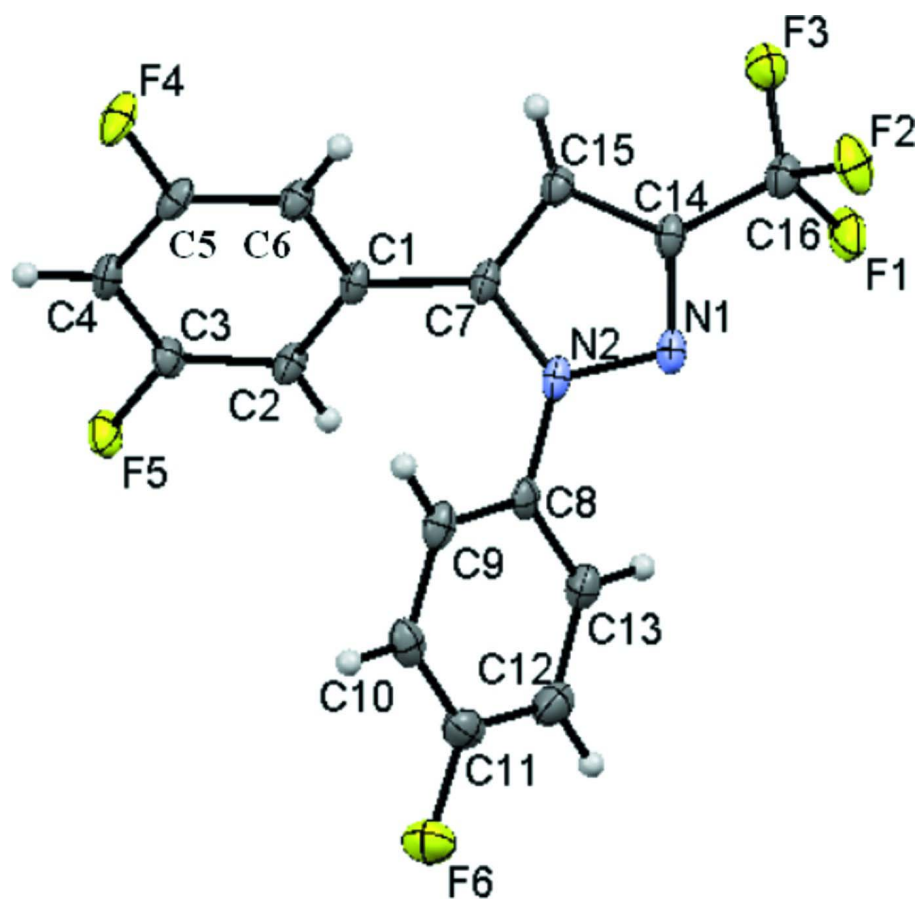
**S2. Experimental**

1-(3,5-Difluorophenyl)-4,4,4-trifluorobutane-1,3-dione (0.015 mmol) and (4-fluorophenyl)hydrazine (0.015 mmol) were taken in dry ethanol (14 ml) and the reaction mixture was stirred for 6 h at 348 K under a nitrogen atmosphere. The reaction was monitored by TLC and the excess solvent was removed by vacuum to obtain the crude compound. It was purified by column chromatography using dichloromethane and ethanol (9:1) as eluent.

Yellow prisms suitable for diffraction studies were grown by slow evaporation of the solvent system: dichloromethane and methanol (9:1).

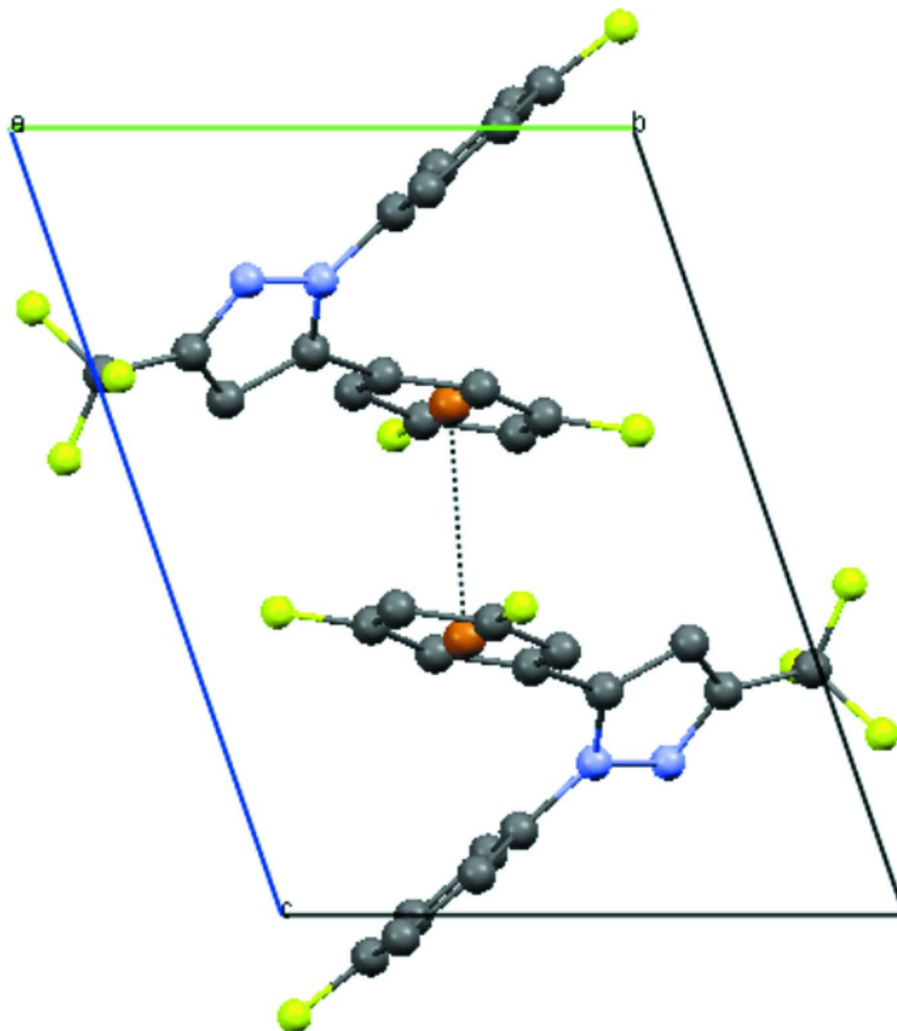
**S3. Refinement**

The H atoms were positioned with idealized geometry using a riding model with C—H = 0.93 Å. The isotropic displacement parameters for all H atoms were set to 1.2 times  $U_{eq}$  of the parent atom.



**Figure 1**

The molecular structure of the title compound, showing the atom-labeling scheme. Displacement ellipsoids are drawn at the 50% probability level.



**Figure 2**

Aromatic  $Cg \cdots Cg$  stacking interactions observed in the crystal structure of the title compound. [ $Cg$  is the centroid of the difluorobenzene ring. H-atoms are omitted for clarity.]

**5-(3,5-Difluorophenyl)-1-(4-fluorophenyl)-3-trifluoromethyl-1*H*-pyrazole**

*Crystal data*

$C_{16}H_8F_6N_2$   
 $M_r = 342.24$   
 Triclinic,  $P\bar{1}$   
 Hall symbol:  $-P\ 1$   
 $a = 7.2535\ (3)\ \text{\AA}$   
 $b = 8.6686\ (4)\ \text{\AA}$   
 $c = 11.7690\ (5)\ \text{\AA}$   
 $\alpha = 70.909\ (1)^\circ$   
 $\beta = 80.139\ (1)^\circ$   
 $\gamma = 88.077\ (1)^\circ$   
 $V = 688.78\ (5)\ \text{\AA}^3$   
 $Z = 2$

$F(000) = 344$   
 Prism  
 $D_x = 1.650\ \text{Mg m}^{-3}$   
 Melting point: 456 K  
 Cu  $K\alpha$  radiation,  $\lambda = 1.54178\ \text{\AA}$   
 Cell parameters from 1234 reflections  
 $\theta = 4.0\text{--}64.8^\circ$   
 $\mu = 1.39\ \text{mm}^{-1}$   
 $T = 293\ \text{K}$   
 Prism, yellow  
 $0.39 \times 0.35 \times 0.29\ \text{mm}$

*Data collection*

Bruker APEXII CCD  
 diffractometer  
 Radiation source: fine-focus sealed tube  
 Graphite monochromator  
 $\varphi$  and  $\omega$  scans  
 Absorption correction: multi-scan  
 (SADABS; Bruker, 2009)  
 $T_{\min} = 0.611$ ,  $T_{\max} = 0.669$

7069 measured reflections  
 2181 independent reflections  
 2040 reflections with  $I > 2\sigma(I)$   
 $R_{\text{int}} = 0.042$   
 $\theta_{\max} = 64.8^\circ$ ,  $\theta_{\min} = 4.0^\circ$   
 $h = -8 \rightarrow 8$   
 $k = -9 \rightarrow 10$   
 $l = -13 \rightarrow 13$

*Refinement*

Refinement on  $F^2$   
 Least-squares matrix: full  
 $R[F^2 > 2\sigma(F^2)] = 0.054$   
 $wR(F^2) = 0.181$   
 $S = 1.11$   
 2181 reflections  
 217 parameters  
 0 restraints  
 Primary atom site location: structure-invariant  
 direct methods

Secondary atom site location: difference Fourier  
 map  
 Hydrogen site location: inferred from  
 neighbouring sites  
 H-atom parameters constrained  
 $w = 1/[\sigma^2(F_o^2) + (0.1212P)^2 + 0.5142P]$   
 where  $P = (F_o^2 + 2F_c^2)/3$   
 $(\Delta/\sigma)_{\max} < 0.001$   
 $\Delta\rho_{\max} = 0.45 \text{ e } \text{\AA}^{-3}$   
 $\Delta\rho_{\min} = -0.48 \text{ e } \text{\AA}^{-3}$

*Special details*

**Geometry.** All s.u.'s (except the s.u. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell s.u.'s are taken into account individually in the estimation of s.u.'s in distances, angles and torsion angles; correlations between s.u.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell s.u.'s is used for estimating s.u.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > 2\sigma(F^2)$  is used only for calculating  $R$ -factors(gt) etc. and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

*Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )*

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
F5	0.4695 (2)	0.16122 (18)	0.61518 (14)	0.0259 (4)
F4	0.0082 (2)	0.5540 (2)	0.60896 (14)	0.0277 (4)
F3	0.8754 (2)	1.09529 (19)	0.57955 (14)	0.0311 (4)
F2	0.8448 (2)	1.0637 (2)	0.77056 (15)	0.0338 (5)
F6	0.7353 (3)	−0.0365 (2)	1.12881 (15)	0.0368 (5)
F1	1.0813 (2)	0.9640 (2)	0.68380 (16)	0.0329 (5)
N2	0.7064 (3)	0.5860 (3)	0.80633 (18)	0.0181 (5)
C8	0.7111 (3)	0.4265 (3)	0.8934 (2)	0.0178 (6)
C2	0.5326 (3)	0.3904 (3)	0.6683 (2)	0.0186 (6)
H2	0.6510	0.3520	0.6815	0.022*
C10	0.5548 (4)	0.1787 (3)	1.0287 (2)	0.0231 (6)
H10	0.4464	0.1152	1.0651	0.028*
C7	0.6008 (3)	0.6408 (3)	0.7159 (2)	0.0178 (5)
C3	0.4110 (4)	0.3037 (3)	0.6324 (2)	0.0188 (5)
C14	0.7893 (3)	0.8337 (3)	0.7149 (2)	0.0191 (6)
C4	0.2343 (3)	0.3540 (3)	0.6116 (2)	0.0196 (6)

H4	0.1552	0.2928	0.5875	0.024*
C5	0.1808 (3)	0.5007 (3)	0.6286 (2)	0.0196 (6)
N1	0.8244 (3)	0.7042 (3)	0.80627 (19)	0.0197 (5)
C16	0.8973 (3)	0.9877 (3)	0.6872 (2)	0.0222 (6)
C11	0.7271 (4)	0.1186 (3)	1.0535 (2)	0.0252 (6)
C6	0.2943 (3)	0.5943 (3)	0.6650 (2)	0.0191 (6)
H6	0.2529	0.6918	0.6763	0.023*
C1	0.4720 (3)	0.5386 (3)	0.6844 (2)	0.0174 (5)
C13	0.8827 (4)	0.3653 (3)	0.9218 (2)	0.0230 (6)
H13	0.9914	0.4288	0.8860	0.028*
C15	0.6500 (3)	0.8044 (3)	0.6551 (2)	0.0187 (6)
H15	0.6018	0.8771	0.5902	0.022*
C9	0.5463 (3)	0.3358 (3)	0.9483 (2)	0.0207 (6)
H9	0.4315	0.3799	0.9314	0.025*
C12	0.8922 (4)	0.2094 (4)	1.0036 (2)	0.0276 (6)
H12	1.0060	0.1672	1.0243	0.033*

*Atomic displacement parameters ( $\text{\AA}^2$ )*

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
F5	0.0287 (8)	0.0245 (8)	0.0340 (9)	0.0044 (6)	−0.0110 (6)	−0.0196 (7)
F4	0.0148 (7)	0.0429 (10)	0.0349 (9)	0.0067 (6)	−0.0103 (6)	−0.0230 (8)
F3	0.0322 (9)	0.0291 (9)	0.0316 (9)	−0.0068 (7)	−0.0084 (7)	−0.0073 (7)
F2	0.0414 (10)	0.0309 (9)	0.0371 (10)	−0.0063 (7)	−0.0009 (7)	−0.0241 (8)
F6	0.0465 (10)	0.0285 (10)	0.0311 (9)	−0.0012 (8)	−0.0096 (8)	−0.0021 (7)
F1	0.0195 (8)	0.0299 (9)	0.0533 (11)	−0.0036 (6)	−0.0088 (7)	−0.0169 (8)
N2	0.0154 (10)	0.0240 (12)	0.0198 (11)	−0.0024 (8)	−0.0037 (8)	−0.0131 (9)
C8	0.0225 (12)	0.0211 (13)	0.0150 (11)	−0.0015 (10)	−0.0050 (9)	−0.0116 (10)
C2	0.0149 (11)	0.0263 (14)	0.0191 (12)	0.0019 (10)	−0.0053 (9)	−0.0125 (10)
C10	0.0284 (13)	0.0241 (14)	0.0184 (13)	−0.0058 (10)	0.0013 (10)	−0.0113 (11)
C7	0.0135 (11)	0.0257 (13)	0.0197 (12)	0.0001 (9)	−0.0044 (9)	−0.0140 (10)
C3	0.0215 (12)	0.0192 (12)	0.0195 (12)	0.0013 (9)	−0.0033 (9)	−0.0116 (10)
C14	0.0180 (12)	0.0240 (13)	0.0197 (12)	−0.0010 (10)	−0.0027 (9)	−0.0135 (11)
C4	0.0178 (12)	0.0267 (14)	0.0180 (12)	−0.0040 (10)	−0.0036 (9)	−0.0114 (10)
C5	0.0115 (11)	0.0315 (14)	0.0173 (12)	0.0010 (10)	−0.0032 (9)	−0.0098 (10)
N1	0.0175 (10)	0.0232 (12)	0.0227 (11)	−0.0031 (8)	−0.0043 (8)	−0.0127 (9)
C16	0.0202 (12)	0.0267 (14)	0.0246 (14)	0.0011 (10)	−0.0054 (10)	−0.0140 (11)
C11	0.0373 (16)	0.0224 (14)	0.0167 (12)	0.0003 (11)	−0.0061 (11)	−0.0067 (10)
C6	0.0182 (12)	0.0235 (13)	0.0194 (12)	0.0025 (10)	−0.0047 (9)	−0.0116 (10)
C1	0.0157 (11)	0.0241 (13)	0.0159 (12)	−0.0018 (9)	−0.0030 (9)	−0.0108 (10)
C13	0.0214 (13)	0.0283 (14)	0.0219 (13)	−0.0032 (10)	−0.0057 (10)	−0.0103 (11)
C15	0.0182 (12)	0.0215 (13)	0.0208 (12)	0.0022 (10)	−0.0059 (9)	−0.0116 (10)
C9	0.0180 (12)	0.0303 (14)	0.0192 (12)	0.0001 (10)	−0.0018 (9)	−0.0157 (11)
C12	0.0278 (14)	0.0321 (15)	0.0258 (14)	0.0027 (11)	−0.0109 (11)	−0.0105 (12)

*Geometric parameters (Å, °)*

F5—C3	1.359 (3)	C7—C15	1.390 (4)
F4—C5	1.351 (3)	C7—C1	1.478 (3)
F3—C16	1.336 (3)	C3—C4	1.376 (4)
F2—C16	1.349 (3)	C14—N1	1.329 (3)
F6—C11	1.353 (3)	C14—C15	1.399 (3)
F1—C16	1.339 (3)	C14—C16	1.483 (3)
N2—N1	1.357 (3)	C4—C5	1.383 (4)
N2—C7	1.367 (3)	C4—H4	0.9300
N2—C8	1.430 (3)	C5—C6	1.383 (3)
C8—C13	1.388 (4)	C11—C12	1.385 (4)
C8—C9	1.390 (4)	C6—C1	1.390 (3)
C2—C3	1.378 (3)	C6—H6	0.9300
C2—C1	1.404 (4)	C13—C12	1.387 (4)
C2—H2	0.9300	C13—H13	0.9300
C10—C11	1.378 (4)	C15—H15	0.9300
C10—C9	1.387 (4)	C9—H9	0.9300
C10—H10	0.9300	C12—H12	0.9300
N1—N2—C7	112.0 (2)	F3—C16—F1	107.1 (2)
N1—N2—C8	118.73 (19)	F3—C16—F2	106.1 (2)
C7—N2—C8	129.3 (2)	F1—C16—F2	106.02 (19)
C13—C8—C9	121.0 (2)	F3—C16—C14	111.8 (2)
C13—C8—N2	118.7 (2)	F1—C16—C14	112.6 (2)
C9—C8—N2	120.3 (2)	F2—C16—C14	112.7 (2)
C3—C2—C1	117.9 (2)	F6—C11—C10	118.5 (2)
C3—C2—H2	121.1	F6—C11—C12	118.7 (2)
C1—C2—H2	121.1	C10—C11—C12	122.8 (3)
C11—C10—C9	118.7 (2)	C5—C6—C1	118.2 (2)
C11—C10—H10	120.6	C5—C6—H6	120.9
C9—C10—H10	120.6	C1—C6—H6	120.9
N2—C7—C15	106.9 (2)	C6—C1—C2	120.5 (2)
N2—C7—C1	125.3 (2)	C6—C1—C7	119.4 (2)
C15—C7—C1	127.6 (2)	C2—C1—C7	120.0 (2)
F5—C3—C4	118.0 (2)	C12—C13—C8	120.0 (2)
F5—C3—C2	118.2 (2)	C12—C13—H13	120.0
C4—C3—C2	123.8 (2)	C8—C13—H13	120.0
N1—C14—C15	113.4 (2)	C7—C15—C14	103.6 (2)
N1—C14—C16	118.8 (2)	C7—C15—H15	128.2
C15—C14—C16	127.8 (2)	C14—C15—H15	128.2
C3—C4—C5	116.2 (2)	C10—C9—C8	119.3 (2)
C3—C4—H4	121.9	C10—C9—H9	120.3
C5—C4—H4	121.9	C8—C9—H9	120.3
F4—C5—C4	117.9 (2)	C11—C12—C13	118.0 (2)
F4—C5—C6	118.7 (2)	C11—C12—H12	121.0
C4—C5—C6	123.4 (2)	C13—C12—H12	121.0
C14—N1—N2	104.07 (19)		

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N1—N2—C8—C13	−37.8 (3)	C9—C10—C11—F6	−177.9 (2)
C7—N2—C8—C13	139.9 (2)	C9—C10—C11—C12	1.4 (4)
N1—N2—C8—C9	142.0 (2)	F4—C5—C6—C1	−179.7 (2)
C7—N2—C8—C9	−40.3 (3)	C4—C5—C6—C1	0.7 (4)
N1—N2—C7—C15	−1.2 (3)	C5—C6—C1—C2	−0.6 (4)
C8—N2—C7—C15	−179.0 (2)	C5—C6—C1—C7	176.2 (2)
N1—N2—C7—C1	174.1 (2)	C3—C2—C1—C6	0.3 (4)
C8—N2—C7—C1	−3.7 (4)	C3—C2—C1—C7	−176.5 (2)
C1—C2—C3—F5	179.2 (2)	N2—C7—C1—C6	134.5 (2)
C1—C2—C3—C4	0.0 (4)	C15—C7—C1—C6	−51.2 (3)
F5—C3—C4—C5	−179.2 (2)	N2—C7—C1—C2	−48.6 (3)
C2—C3—C4—C5	0.0 (4)	C15—C7—C1—C2	125.7 (3)
C3—C4—C5—F4	180.0 (2)	C9—C8—C13—C12	1.8 (4)
C3—C4—C5—C6	−0.4 (4)	N2—C8—C13—C12	−178.4 (2)
C15—C14—N1—N2	0.5 (3)	N2—C7—C15—C14	1.4 (2)
C16—C14—N1—N2	−179.3 (2)	C1—C7—C15—C14	−173.8 (2)
C7—N2—N1—C14	0.5 (2)	N1—C14—C15—C7	−1.2 (3)
C8—N2—N1—C14	178.55 (19)	C16—C14—C15—C7	178.5 (2)
N1—C14—C16—F3	167.2 (2)	C11—C10—C9—C8	1.3 (3)
C15—C14—C16—F3	−12.5 (3)	C13—C8—C9—C10	−2.9 (3)
N1—C14—C16—F1	46.5 (3)	N2—C8—C9—C10	177.3 (2)
C15—C14—C16—F1	−133.1 (3)	F6—C11—C12—C13	176.8 (2)
N1—C14—C16—F2	−73.3 (3)	C10—C11—C12—C13	−2.4 (4)
C15—C14—C16—F2	107.0 (3)	C8—C13—C12—C11	0.8 (4)

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